



# Nickel-Catcher-Doped Zwitterionic Hydrogel Coating on Nickel–Titanium Alloy Toward Capture and Detection of Nickel Ions

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### Specialty section:

This article was submitted to  
Biosensors and Biomolecular  
Electronics,  
a section of the journal  
Frontiers in Bioengineering and  
Biotechnology

**Received:** 22 April 2021

**Accepted:** 25 May 2021

**Published:** 24 June 2021

### Citation:

Fu X, Liu X, Hao D, Xiao W, Nie Q  
and Meng J (2021)  
Nickel-Catcher-Doped Zwitterionic  
Hydrogel Coating on Nickel–Titanium  
Alloy Toward Capture and Detection  
of Nickel Ions.  
Front. Bioeng. Biotechnol. 9:698745.  
doi: 10.3389/fbioe.2021.698745

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Nickel–titanium (NiTi) alloys show broad applicability in biomedical fields. However, the unexpected aggregation of bacteria and the corrosion of body fluid on NiTi-based medical devices often lead to the leakage of nickel ions, resulting in inevitable allergic and cytotoxic activities. Therefore, the capture and detection of nickel ions are important to avoid serious adverse reactions caused by NiTi-based medical devices. Herein, we presented a nickel ion capture strategy by the combination of zwitterionic hydrogels as anti-bacteria layers and carbon disulfide (CS<sub>2</sub>) components as nickel-catchers (Ni-catchers). On the one hand, the hydration layer of zwitterionic hydrogel can efficiently inhibit bacteria adhesion and reduce nickel ions leakage from NiTi corrosion. On the other hand, Ni-catchers can capture leaked nickel ions from NiTi alloy actively by chelation reaction. Therefore, this strategy shows great capabilities in resisting bacteria adhesion and capturing nickel ions, providing the potential possibility for the detection of nickel ion leakage for implantable biomedical materials and devices.

**Keywords:** nickel ion capture, anti-bacteria adhesion, zwitterionic hydrogel, surface modification, nickel–titanium alloy

## INTRODUCTION

With the rapid development of biomedicine, implantable materials have been widely used in diagnosis, treatment, and prosthesis (Sylvestre et al., 2020). As one of the most popular biomaterials, nickel–titanium (NiTi) alloys have been applied in orthodontic archwires (Liu et al., 2019), bone implants (Raphel et al., 2016), and vascular stents (Robertson and Ritchie, 2007), due to their unique superelasticity and good biocompatibility. However, the leakage of nickel ions (Ni<sup>2+</sup> ions) from NiTi alloys often leads to adverse reactions such as allergic signs (e.g., contact mucositis and eczematous and urticarial reactions; Huang et al., 2003; Kolokitha and Chatzistavrou, 2008) and cytotoxic responses (Jia et al., 1999). In addition, the unexpected adhesion of bacteria and acidic oral environment corrode NiTi alloys, aggravating the leakage of Ni<sup>2+</sup> ions (Huang et al., 2003). Therefore, a novel strategy for capturing and detecting leaked Ni<sup>2+</sup> ions is urgently needed for potential biomedical applications.

In recent years, many passive strategies have been developed to avoid the leakage of  $\text{Ni}^{2+}$  ions, including growing oxide films [e.g.,  $\text{TiO}_2$  layer (Gu et al., 2005) and  $\text{NiTiO}_3$  nanosheet (Hang et al., 2019)], depositing coatings [e.g., diamond-like carbon coating (Hang et al., 2012) and 1H,1H,2H,2H-perfluorodecyltrimethoxysilane (FAS) deposited layer (Liu et al., 2019)], and implanting barrier elements [e.g., carbon (Poon et al., 2005) and oxygen (Wu et al., 2006) plasma immersion ion implantation]. However, most of them can not prohibit long-term leakage of  $\text{Ni}^{2+}$  ions because of the low stability of hydrophobic surface in acidic oral environment (Liu et al., 2019) and the corrosion of adhered bacteria (Gu et al., 2005). Moreover, there are many anti-bacteria methods including zwitterionic hydrogels (Li et al., 2008; Shivapooja et al., 2015; Lin et al., 2020), polyethylene glycol (Peng et al., 2017), quaternary ammonium salt groups (Wei et al., 2016, 2017), protein-based surfaces (Liu et al., 2018), and multiwalled carbon nanotubes (Hartono et al., 2018). Thus, combining anti-bacteria and  $\text{Ni}^{2+}$  ion capture may be a promising strategy for the new generation of NiTi-based medical devices. As a conventional medical treatment for acute carbonyl nickel poisoning, carbon disulfide ( $\text{CS}_2$ ) derivatives can serve as Ni-catchers by chelation reaction (Sunderman, 1990). Therefore, we believe that introducing Ni-catchers into anti-bacteria layer can achieve long-term  $\text{Ni}^{2+}$  ion capture and detection.

Herein, we present a Ni-catcher-doped zwitterionic hydrogel coating on the surface of NiTi alloy, showing robust anti-bacteria and long-term  $\text{Ni}^{2+}$  ion capture (Figure 1). In brief, zwitterionic hydrogel was firstly prepared on the surface of NiTi alloy (H-NiTi) by the copolymerization of 2-hydroxyethyl methacrylate (HEMA), 3-dimethyl(methacryloyloxyethyl)ammonium propane-sulfonate (DMAPS), and methyl methacrylate (MMA; Su et al., 2020). Then, the  $\text{CS}_2$  as a Ni-catcher was grafted into zwitterionic hydrogel by immersing H-NiTi in the mixture of  $\text{CS}_2$  and triethylamine (TEA) to obtain CS-NiTi (Figure 1A and Supplementary Figure 1). The zwitterionic hydrogel can prevent the adhesion of bacteria by the hydration layer on the surface of CS-NiTi. Moreover, the Ni-catcher endows this coating with  $\text{Ni}^{2+}$  ion capture through chelation reaction. Therefore, this work provides a promising avenue to capture and detect  $\text{Ni}^{2+}$  ions, which could be used for implantable biomedical materials and devices.

## MATERIALS AND METHODS

### Chemicals and Materials

Nickel–titanium alloy was purchased from Sunway Technology Co. Ltd. (Guangdong, China). Brain heart infusion (BHI) was obtained from Oxoid Co. (Hampshire, United Kingdom). Dulbecco's modified Eagle medium (DMEM) was bought from Thermo Fisher Biochemical Products Co. Ltd. (Beijing, China). Tris–HCl buffer was purchased from Coolaber Science Technology (Beijing, China). Phosphate-buffered saline (PBS) was bought from HyClone Laboratories Inc. (Logan, UT, United States). Acetone (>99.5%, AR) and alcohol ( $\geq 99.8\%$ , GR) were obtained from the Beijing Chemical Co. (Beijing,

China). Bis(2-methacryloxyethyl)phosphate (bis-HEMA) and 2,2-diethoxyacetophenone (DEAP) were bought from J&K Scientific (Beijing, China). HEMA and DMAPS were purchased from Aladdin Ltd. (Shanghai, China). MMA and poly(ethylene glycol)dimethacrylate (PEGDMA) were purchased from Alfa Aesar (Haverhill, MA, United States).  $\text{CS}_2$  and TEA were obtained from Macklin Inc. (Shanghai, China). Deionized water (>1.82 M $\Omega$  cm, Milli-Q system, Merck & Co., Kenilworth, NJ, United States) was used. All reagents were used as received.

### Preparation of CS-NiTi

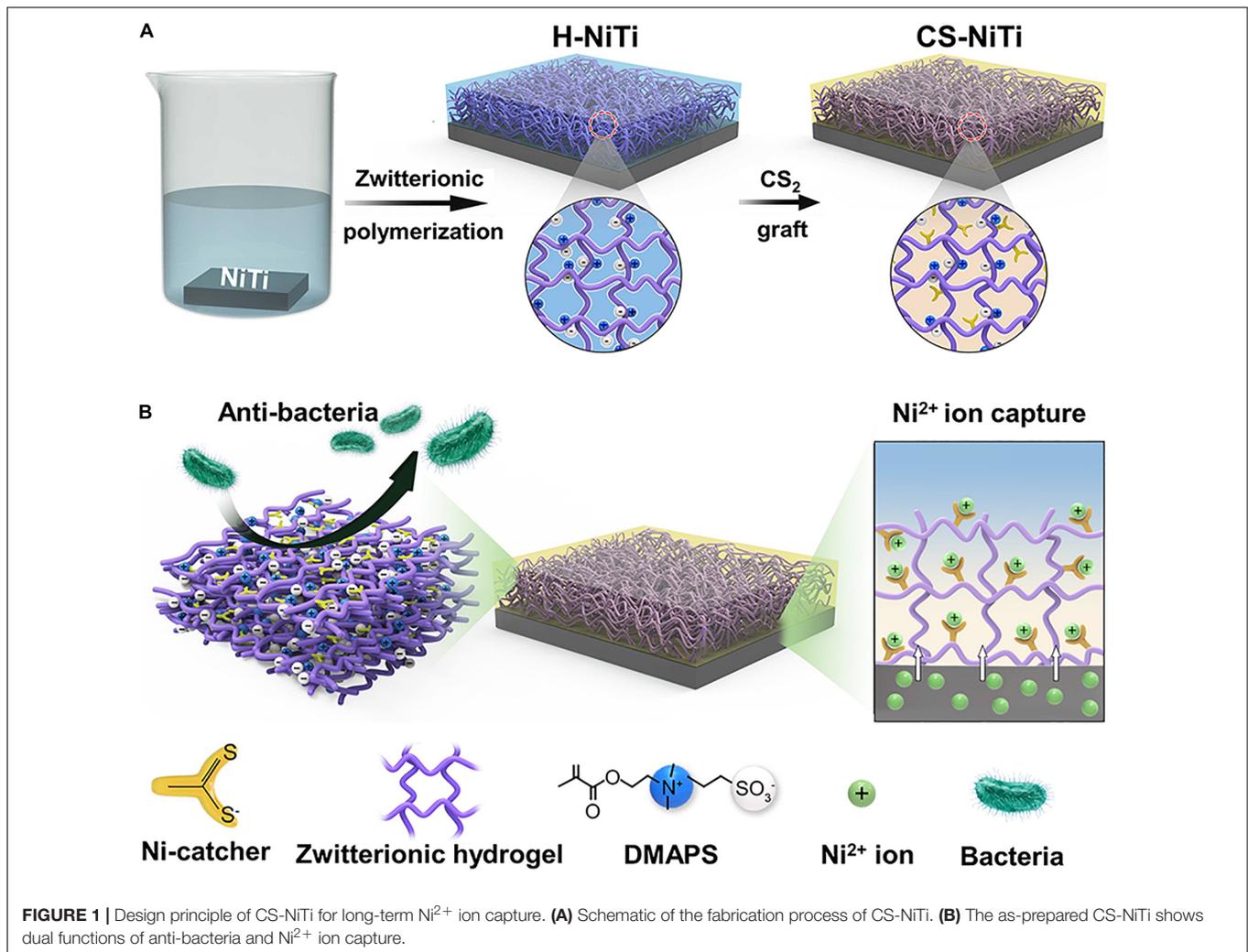
The CS-NiTi was fabricated according to previous literature, with some modifications (Su et al., 2020). Firstly, pristine NiTi alloy was immersed in bis-HEMA (5% ethanol solution) for 24 h and washed with deionized water for three times. Then, the 0.5 wt% of the photoinitiator DEAP was coated on the surface of NiTi alloy. The hydrogel precursor solution was obtained by mixing the crosslinker of PEGDMA (2 wt%) with the hydrogel monomers solution, including DMAPS (5 wt%), HEMA (40 wt%), and MMA (5 wt%). Then, it was degassed for 20 min by KQ5200DE ultrasonic cleaner (Kun Shan Ultrasonic Instruments Co. Ltd., Shanghai, China). The degassed precursor solution was poured on the surface of NiTi alloy and exposed to UV irradiation (wavelength  $\approx 365$  nm) in a dark chamber for 600 s by a UV LED curing system (Beijing NBET Technology Co. Ltd., Beijing, China), to obtain H-NiTi. To graft Ni-catchers on H-NiTi, H-NiTi was immersed into a mixture for 24 h containing  $\text{CS}_2$ , TEA, and deionized water. LCS-NiTi, MCS-NiTi, and HCS-NiTi were obtained by regulating the concentration of  $\text{CS}_2$  (0.2, 0.6, and 1.0 mol/L), respectively.

### Characterizations

The morphology of pristine NiTi, H-NiTi, and CS-NiTi were observed with S-4800 scanning electron microscope (SEM; Hitachi, Tokyo, Japan) and Single Lens Reflex camera (SLR camera; Nikon D80, Tokyo, Japan). Water contact angles (WCAs) were measured by dropping 2  $\mu\text{l}$  of water on the surface of pristine NiTi, H-NiTi, and CS-NiTi by the Dataphysics OCA 20 Contact-Angle System (Filderstadt, Germany). The S element analysis was performed on ESCALAB 250Xi X-ray photoelectron spectroscopy analysis (XPS; Thermo Fisher Scientific, Waltham, MA, United States). The transmittance was measured by UV-2600 UV-VIS spectrophotometer (SHIMADZU, Kyoto, Japan) with wavelength from 450 to 700 nm.

### Cell Culture and Cellular Toxicity Test

Cell Counting Kit-8 (CCK-8) assay was performed to investigate the cell viability (Xiao et al., 2020). And the cytotoxicity of materials including pristine NiTi, H-NiTi, and CS-NiTi was evaluated with HepG2 cells. HepG2 cells with the density of  $1 \times 10^5$  cells per well were seeded into the 24-well plates in DMEM containing 10% fetal bovine serum (FBS) and 1% antibiotic solution (penicillin and streptomycin) and then cultured at 37 C in a humidified atmosphere with 5%  $\text{CO}_2$  for 36 h. Pristine NiTi, H-NiTi, and CS-NiTi were added to each well, and the cells were incubated for 12 h. Subsequently, CCK-8 solutions were added to the wells followed by 2-h incubation at



37°C. And then the supernatant was obtained and added to 96-well plates. Finally, microplate reader was used to measure the absorbance values per well at a test wavelength of 450 nm and a reference wavelength of 690 nm to analyze the cell viability. Control experiments were done under the same condition with PBS. Eq. (1) was used to calculate the cell viability rate:

$$\text{Cell viability} = A_s/A_c \times 100\% \quad (1)$$

where  $A_s$  represents the absorbance of the treatments of different groups and  $A_c$  represents the absorbance of the control treatments with PBS. All the experiments were performed in triplicate.

## Bacteria Cultivation

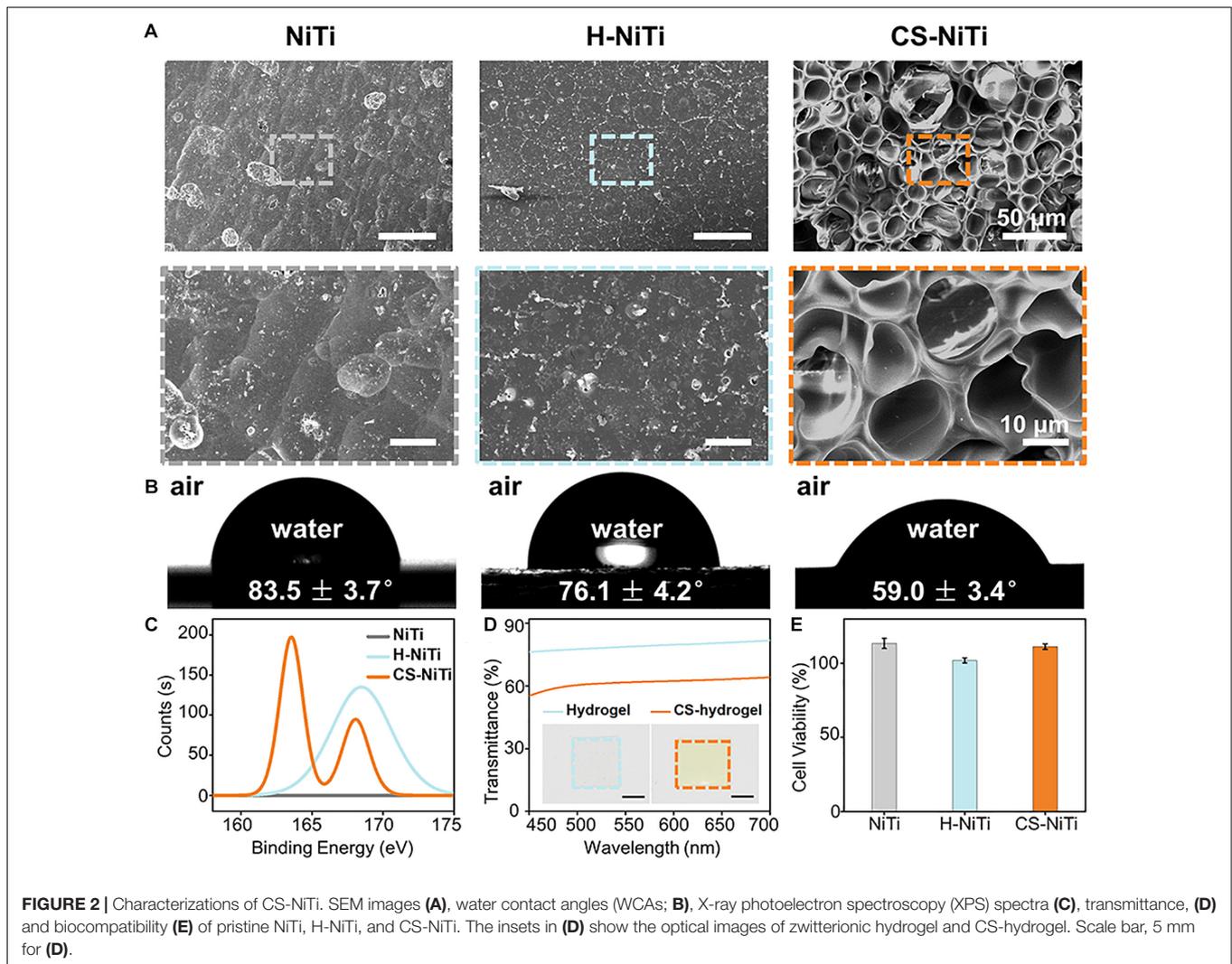
*Streptococcus mutans* (UA159) was used as the model bacteria. *S. mutans* was revived by incubating its freezing solution with 1 ml of BHI at 37°C for 48 h. Then the suspension of *S. mutans* was centrifuged at 2,000 rpm for 10 min at room temperature to discard the supernatant. Then, the *S. mutans* aggregation was resuspended in another 1 ml of BHI and incubation at 37°C for another 48 h. After the logarithmic phase was reached, the

bacteria were harvested by centrifugation at 2,000 rpm for 10 min at room temperature and washed with Tris-HCl buffer three times. Subsequently, the bacteria were suspended in 50 ml of Tris-HCl buffer at a final concentration of  $1 \times 10^8$  CFU/ml (Liu et al., 2018).

## Anti-bacteria Property

To test anti-bacteria ability, the pristine NiTi, H-NiTi, and CS-NiTi were incubated in the as-prepared bacteria suspensions for 1, 3, 6, 12, 18, and 24 h at 37°C. After the incubation, they were taken out and fixed in 2.5% glutaraldehyde for 1 h at room temperature, followed by immersing in a graded series of alcohol concentrations (30, 50, 70, 80, 90, and 100%) for 10 min to dehydrate the adhered bacteria. SEM was used to observe bacteria adhesion performance. The quantitative evaluation of the adhered bacteria was obtained by Cell Profiler 4.0.7. The anti-bacteria efficiency ( $E_{\text{anti-bacteria}}$ ) could be calculated by Eq. (2):

$$E_{\text{anti-bacteria}} = (1 - D_{\text{CS-NiTi}}/D_{\text{NiTi}}) \times 100\% \quad (2)$$



where  $D_{CS-NiTi}$  and  $D_{NiTi}$  refer to the bacteria density on the surface of CS-NiTi and pristine NiTi, respectively.

### Performance of Ni<sup>2+</sup> Ion Capture

The performance of Ni<sup>2+</sup> ion capture was investigated on 710-OES inductively coupled plasma emission spectrometer (ICP; Varian, Palo Alto, CA, United States) by detecting the concentration of nickel ions in the supernatant of the remaining bacteria suspension incubated with pristine NiTi, H-NiTi, LCS-NiTi, MCS-NiTi, and HCS-NiTi for different times at 37°C. The capture efficiency of Ni<sup>2+</sup> ions ( $E_{capture}$ ) could be calculated by Eq. (3):

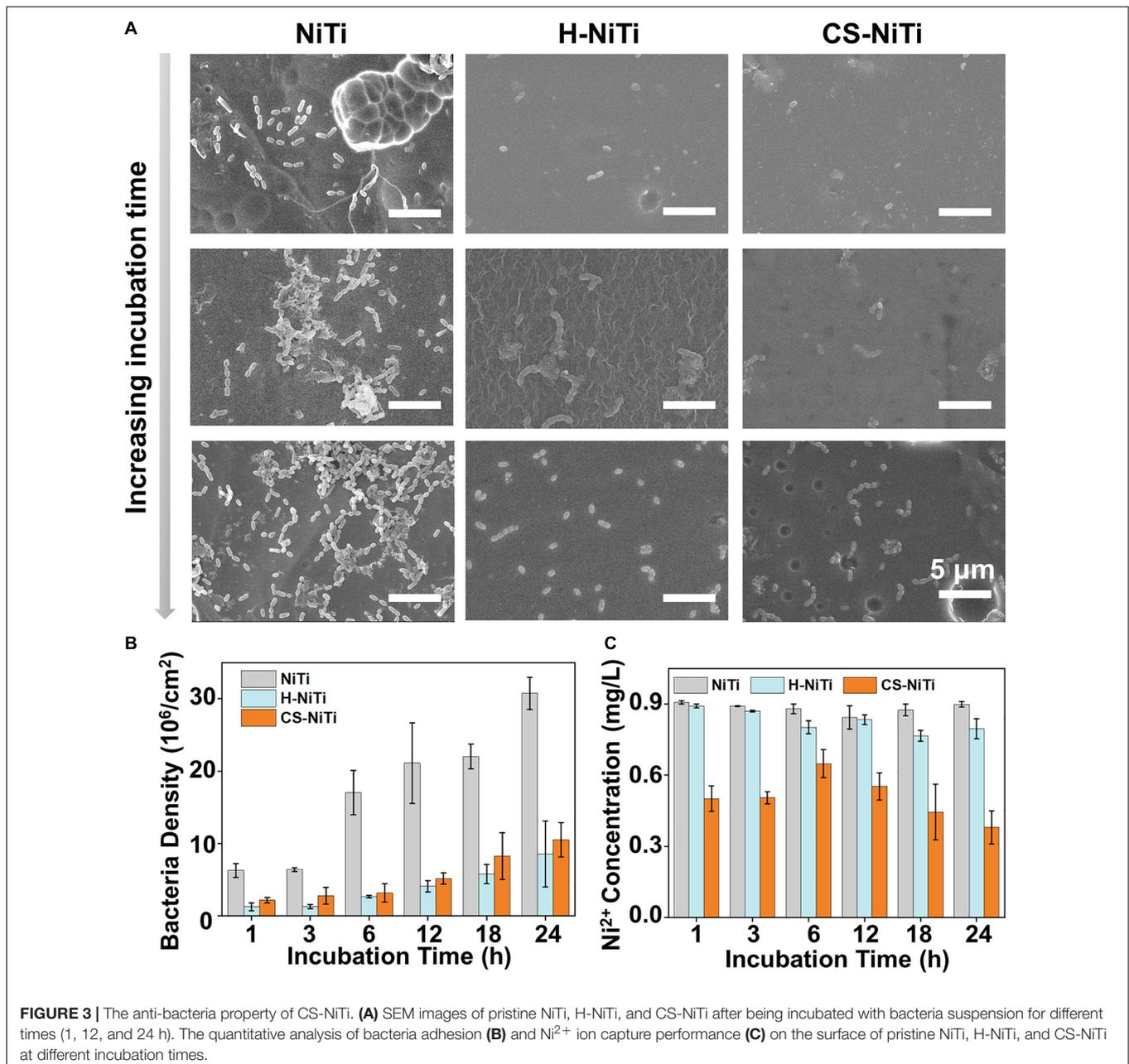
$$E_{capture} = (1 - C_{CS-NiTi}/C_{H-NiTi}) \times 100\% \quad (3)$$

where  $C_{CS-NiTi}$  and  $C_{H-NiTi}$  refer to the concentration of Ni<sup>2+</sup> ions in the supernatant of remaining bacteria suspension incubated with CS-NiTi and H-NiTi, respectively.

## RESULTS

### Characterizations of CS-NiTi

To verify the introduction of Ni-catchers, we compared the pristine NiTi, H-NiTi, and CS-NiTi from surface morphology, WCAs, XPS spectra, transmittance, optical images, and biocompatibility. As shown in **Figure 2A**, the introduction of Ni-catchers resulted in the porous morphology of CS-NiTi, compared with the flat morphology of H-NiTi, probably because of the phase separation (Su et al., 2020). Furthermore, the WCAs decreased from  $83.5 \pm 3.7^\circ$  (NiTi) to  $59.0 \pm 3.4^\circ$  (CS-NiTi), indicating the hydrophilicity of CS-NiTi (**Figure 2B**). The change of wettability might be attributed to the formation of hydrophilic xanthates in the presence of TEA (Weißl et al., 2018) and the porous morphology of CS-NiTi. More hydrophilic surface was beneficial to the formation of the hydration layer, enhancing the anti-bacteria adhesion ability of NiTi alloys (Lin et al., 2020). XPS was further employed to verify the presence of Ni-catchers in CS-NiTi. As shown in **Figure 2C** and **Supplementary Figure 2**, two S 2p peaks at 163 and 168 eV were observed in the XPS spectrum



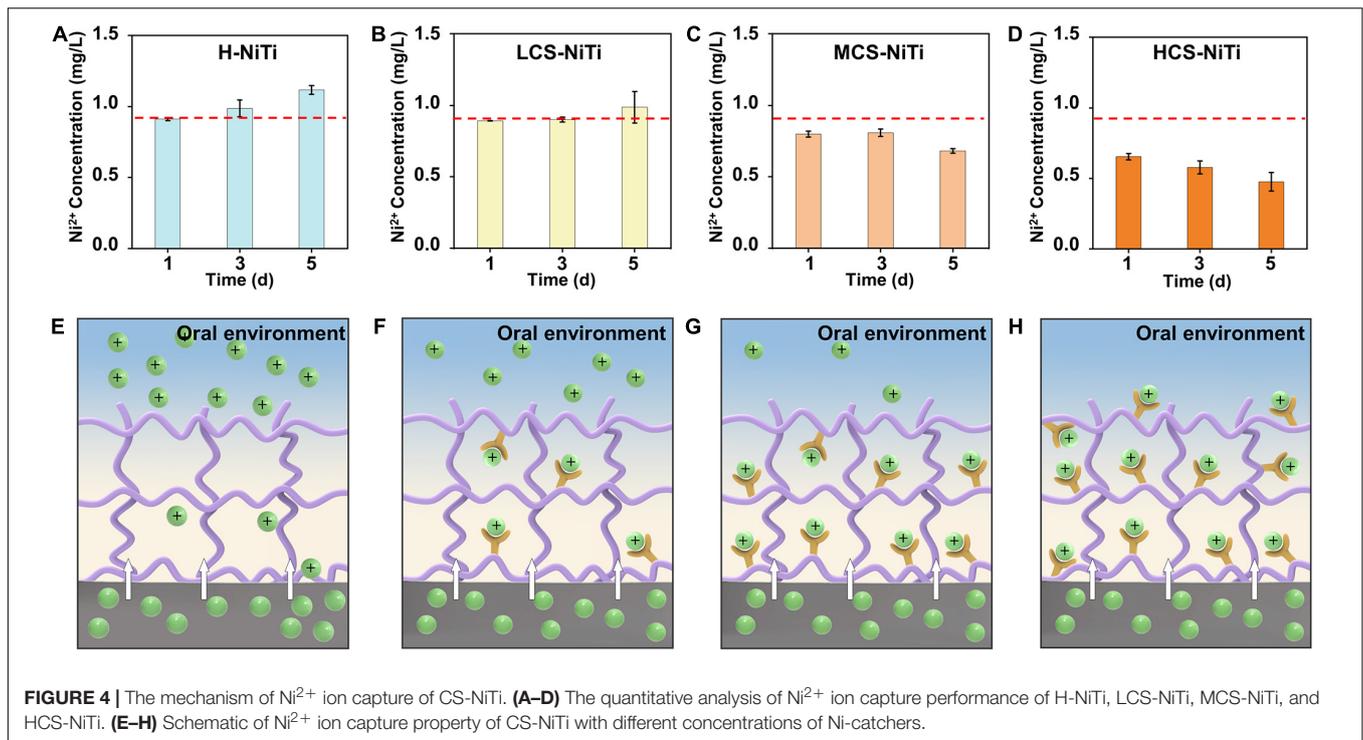
**FIGURE 3 |** The anti-bacteria property of CS-NiTi. **(A)** SEM images of pristine NiTi, H-NiTi, and CS-NiTi after being incubated with bacteria suspension for different times (1, 12, and 24 h). The quantitative analysis of bacteria adhesion **(B)** and Ni<sup>2+</sup> ion capture performance **(C)** on the surface of pristine NiTi, H-NiTi, and CS-NiTi at different incubation times.

of CS-NiTi, compared with the single S 2p peak at 168 eV for H-NiTi and no S 2p peak for NiTi. The S 2p peak at 168 eV shown in H-NiTi represented the S element in zwitterionic hydrogel network. The new S 2p peak at 163 eV indicated the successful introduction of Ni-catchers in CS-NiTi (Deng et al., 2013). Moreover, the appearance of zwitterionic hydrogels changed from colorless to yellow with lower transmittance after Ni-catcher (CS<sub>2</sub>) graft (Figure 2D), for xanthates always appear as a series of yellow compounds. The transmittance of Ni-catcher-doped zwitterionic hydrogel coating was higher than 60%, which has an insignificant impact on the appearance of NiTi alloys and would not affect the application of NiTi alloys in the implantable biomedical field. Furthermore, we tested the

biocompatibility of CS-NiTi using CCK-8 test (Xiao et al., 2020). Higher HepG2 cell viability in Figure 2E indicated the excellent biocompatibility of CS-NiTi, showing promising applicability in biomedical field. Therefore, we successfully fabricated CS-NiTi with good biocompatibility.

### Anti-bacteria Property of CS-NiTi

As a kind of widely used orthodontic material, NiTi alloys are often exposed to the oral environment. However, cleaning teeth is difficult during orthodontic treatment, leading to the adhesion of bacteria and causing caries and periodontitis. To investigate the anti-bacteria property in the oral environment, CS-NiTi was incubated in  $1 \times 10^8$  CFU/ml bacteria suspension (*S. mutans*, the



largest proportion of bacteria in oral environment) for different times, taking pristine NiTi and H-NiTi as controls. As shown in **Figure 3A** and **Supplementary Figure 3**, many bacteria adhered on the surface of pristine NiTi after being incubated for 1 h. In contrast, there were a few bacteria on the surfaces of H-NiTi and CS-NiTi, indicating their good anti-bacteria properties. By prolonging incubation time, a growing number of bacteria adhered on the surface of pristine NiTi and formed a biofilm at 12 h. In contrast, the bacteria density was still very low on the surfaces of H-NiTi and CS-NiTi. The quantitative analysis in **Figure 3B** revealed a good anti-bacteria efficiency in H-NiTi and CS-NiTi at 6 h, showing the highest efficiency of ca. 84% for H-NiTi and ca. 81% for CS-NiTi. By prolonging the incubation time to 24 h, the anti-bacteria efficiency maintained higher than 72% for H-NiTi and 66% for CS-NiTi. These results indicated that CS-NiTi had excellent long-term anti-bacteria property. Moreover, ICP was employed for exploring the performance of  $\text{Ni}^{2+}$  ion capture by monitoring the concentration of  $\text{Ni}^{2+}$  ions in bacteria suspension ( $1 \times 10^8$  CFU/ml). As shown in **Figure 3C**, the concentration of  $\text{Ni}^{2+}$  ions in CS-NiTi was significantly lower than that of NiTi and H-NiTi. Even though it was incubated for 24 h, the concentration of  $\text{Ni}^{2+}$  ions for CS-NiTi was still as low as  $0.38 \pm 0.07$  mg/L, indicating the excellent ability of  $\text{Ni}^{2+}$  ion capture of CS-NiTi. Therefore, CS-NiTi showed dual properties of anti-bacteria and  $\text{Ni}^{2+}$  ion capture at the same time.

## Mechanism of $\text{Ni}^{2+}$ Ion Capture

To further investigate the performance of  $\text{Ni}^{2+}$  ion capture, a series of CS-NiTi were fabricated by regulating the  $\text{CS}_2$  concentration in the mixed solution, that is, 0.2 mol/L of  $\text{CS}_2$  for low concentration (denoted as LCS-NiTi), 0.6 mol/L of  $\text{CS}_2$

for medium concentration (MCS-NiTi), and 1.0 mol/L of  $\text{CS}_2$  for high concentration (HCS-NiTi). With the increasing concentration of Ni-catchers, the surface became more hydrophilic, revealing by the WCAs changing from  $73.0 \pm 2.8^\circ$  to  $59.0 \pm 3.4^\circ$ , along with the increased pore size (**Supplementary Figures 4A,B**). Furthermore, the appearance of CS-hydrogel changed from colorless with high transparency to yellow with reduced transparency (**Supplementary Figure 4C**). To verify the long-term ability of  $\text{Ni}^{2+}$  ion capture, we incubated H-NiTi, LCS-NiTi, MCS-NiTi, and HCS-NiTi with *S. mutans* suspensions for different days. After the incubation, the concentration of leaked  $\text{Ni}^{2+}$  ions from NiTi alloy was reduced from  $0.91 \pm 0.01$  mg/L (H-NiTi) to  $0.65 \pm 0.02$  mg/L (HCS-NiTi) with the incubation time of 1 day (**Figures 4A–D**). By prolonging the incubation time to 5 days, the concentration of leaked  $\text{Ni}^{2+}$  ions increased to  $1.12 \pm 0.03$  mg/L for H-NiTi gradually. On the contrary, the concentration of leaked  $\text{Ni}^{2+}$  ions in HCS-NiTi reduced from  $0.65 \pm 0.02$  to  $0.48 \pm 0.07$  mg/L dramatically, indicating its excellent long-term property of  $\text{Ni}^{2+}$  ion capture.

The possible mechanism of  $\text{Ni}^{2+}$  ion capture is proposed in **Figures 4E–H**. Due to the corrosion of the adhered bacteria and acidic oral microenvironment,  $\text{Ni}^{2+}$  ions are released slowly and persistently from NiTi alloy (Huang et al., 2003). Briefly, the leaked  $\text{Ni}^{2+}$  ions can pass through the zwitterionic hydrogel coating on the surface of NiTi alloy and be released to oral environment easily (**Figure 4E**), as the diffusion of  $\text{Ni}^{2+}$  ions is not efficiently restricted in single zwitterionic hydrogel. In contrast, the concentration of  $\text{Ni}^{2+}$  ions is dramatically decreased after introducing Ni-catchers because grafted  $\text{CS}_2$  can capture free  $\text{Ni}^{2+}$  ions and restrain their diffusion to oral environment (**Figures 4F–H**). Moreover, the long-term capture efficiency of

Ni<sup>2+</sup> ions is improved with the increasing concentration of Ni-catchers. Therefore, the introduction of Ni-catchers endows CS-NiTi with robust Ni<sup>2+</sup> ion capture property.

## DISCUSSION

In summary, we demonstrate a long-term Ni<sup>2+</sup> ion capture coating on the surface of NiTi alloy. The zwitterionic hydrogels and Ni-catchers endow CS-NiTi with excellent anti-bacteria and Ni<sup>2+</sup> ion capture properties. The CS-NiTi exhibited long-term anti-bacteria efficiency of ca. 66% at 24 h and Ni<sup>2+</sup> ion capture efficiency of ca. 57% at 5 days. Furthermore, by integrating chelation reaction and introducing color change, this coating may apply to Ni<sup>2+</sup> ion capture and detection. Therefore, this coating provides a new clue to design the next generation of biomaterials and devices with capture and detection of Ni<sup>2+</sup> ions.

## DATA AVAILABILITY STATEMENT

The raw data supporting the conclusions of this article will be made available by the authors, without undue reservation.

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## AUTHOR CONTRIBUTIONS

QN and JM conceived the idea. DH and XF contributed the fabrication process. XF and WX performed the experiments. XF and XL analyzed the results. XF, XL, QN, and JM wrote the manuscript. All authors contributed to the article and approved the submitted version.

## FUNDING

This research is supported by the National Natural Science Foundation of China (Nos. 51973003, 21875269, and 21988102), National Key R&D Program of China (Nos. 2019YFA0709300 and 2018YFC1105301), and Youth Innovation Promotion Association, CAS (No. 2017036).

## SUPPLEMENTARY MATERIAL

The Supplementary Material for this article can be found online at: <https://www.frontiersin.org/articles/10.3389/fbioe.2021.698745/full#supplementary-material>

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**Conflict of Interest:** The authors declare that the research was conducted in the absence of any commercial or financial relationships that could be construed as a potential conflict of interest.

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